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(Mole fraction is a concentration unit

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introduced in the chapter on gases.)  
Recalling that the total pressure of a gaseous mixture is equal to the sum of partial pressures for all its components (Dalton's law of partial pressures), the total vapor pressure exerted by a solution containing  $i$  components is

### **11.4 Colligative Properties -**



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One mole of glycine,  $C_2H_5O_2N$ , contains 2 moles of carbon, 5 moles of hydrogen, 2 moles of oxygen, and 1 mole of nitrogen: The provided mass of glycine ( $\sim 28$  g) is a bit more than one-third the molar mass ( $\sim 75$  g/mol), so we would expect the computed result to be a bit greater than one-third of a mole ( $\sim 0.33$  mol).

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## **3.1 Formula Mass and the Mole Concept - Chemistry**

What is a Mole? In the field of chemistry, a mole is defined as the amount of a substance that contains exactly  $6.02214076 \times 10^{23}$  'elementary entities' of the given substance.. The number  $6.02214076 \times 10^{23}$  is popularly

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known as the Avogadro constant and is often denoted by the symbol 'N A '. The elementary entities that can be represented in moles can be atoms, molecules, monoatomic ...

### **Mole Concept - What is a Mole? [Related Formulae, Examples]**

The number of moles of CO<sub>2</sub> when 1

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mole of ethane is burned =  $4/2 = 2$   
mole. The number of moles when 10  
moles of ethane burns =  $2 \times 10 = 20$   
mole. Weight of 20 moles =  $20 \times 44 =$   
880 g

**Kerala Syllabus 10th Standard  
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## **chemmybear.com - Resources for Chemistry and AP Chemistry!**

NCERT Solutions Class 11 Chemistry Chapter 2 Structure Of Atom. Here on AglaSem Schools, you can access to NCERT Book Solutions in free pdf for Chemistry for Class 11 so that you can refer them as and when required. The NCERT Solutions to the questions after

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every unit of NCERT textbooks aimed at helping students solving difficult questions.. For a better understanding of this chapter, you ...

### **NCERT Solutions for Class 11 Chemistry Chapter 2 Structure ...**

In Chapter 6, you were introduced to the polymers of life and their building block

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structures, as shown below in Figure 11.1. Recall that the monomer units for building the nucleic acids, DNA and RNA, are the nucleotide bases, whereas the monomers for proteins are amino acids, for carbohydrates are sugar residues, and for lipids are fatty acids ...

### **CH103 - Chapter 8: The Major**

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## **Macromolecules - Chemistry**

NCERT Solutions Class 11 Chemistry  
Chemistry Lab Manual Chemistry Sample  
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SOLVED. Question 1. Calculate the  
molecular mass of the following: (i) H<sub>2</sub>O  
(ii) CO<sub>2</sub> (iii) CH<sub>4</sub> Answer: (i) Molecular  
mass of H<sub>2</sub>O = 2(1.008 amu) + 16.00  
amu = 18.016 amu (ii) Molecular mass of

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$\text{CO}_2 = 12.01 \text{ amu} + 2 \times 16.00 \text{ amu} = 44.01 \text{ amu}$  (iii) Molecular mass of  $\text{CH}_4 = 12.01 \text{ amu} + 4 (1.008 \dots)$

### **NCERT Solutions for Class 11 Chemistry Chapter 1**

You can learn the mole concept in detail in Atoms and Molecules Class 9. Exercise 3.5 total Solutions: 11 Questions (6 short

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question and 5 Long questions). Key Features of NCERT Solutions for Class 9 Science Chapter 3. NCERT Solutions for Class 9 Science Chapter 3 has the following features. Chapter 3 Science Class 9, is presented comprehensively.

**NCERT Solutions for Class 9 Science Chapter 3 Atoms and ...**

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(b) 0.1 mole of  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  has mass =  $322 \times 0.1 = 32.2 \text{ g}$   
(c) 0.1 mole of  $\text{CaCl}_2$  has mass =  $111 \times 0.1 = 11.1 \text{ g}$   
(d) 0.1 mole of Mg has mass =  $24 \times 0.1 = 2.4 \text{ g}$ .  
Solution 14. 1 molecule of  $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$  contains oxygen atoms = 13  
So,  $6.023 \times 10^{23}$  molecules (1 mole) has atoms =  $13 \times 6.023 \times 10^{23}$   
So, 0.1 mole will have atoms =  $0.1 \times 13 \times 6.023$



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...

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Solving the Atoms and Molecules

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help you to revise complete Syllabus and boost your score more in examinations.

## **NCERT Solutions for Class 11 Chemistry Chapter 1 Some ...**

(b) One mole of oxygen gas contains Avogadro's number of molecules. (c) One mole of hydrogen gas contains Avogadro's number of atoms. (d) One

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mole of electrons stands for  $6.023 \times 10^{23}$  electrons. Answer: (a) One gram of C - 12 contains Avogadro's number of atoms. Hint: 12 g of Carbon contains  $6.023 \times 10^{23}$  atoms,

### **Samacheer Kalvi 10th Science Guide Chapter 7 Atoms and ...**

Note: In the following questions two or

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more options may be correct. (Q.11 to Q.14) 11. The positive value of the standard electrode potential of  $\text{Cu}^{2+}/\text{Cu}$  indicates that \_\_\_\_\_ [NCERT Exemplar] (a) this redox couple is a stronger reducing agent than the  $\text{H}^{+}/\text{H}_2$  couple. (b) this redox couple is a stronger oxidising agent than  $\text{H}^{+}/\text{H}_2$ . (c) Cu can displace  $\text{H}_2$  from acid.

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